

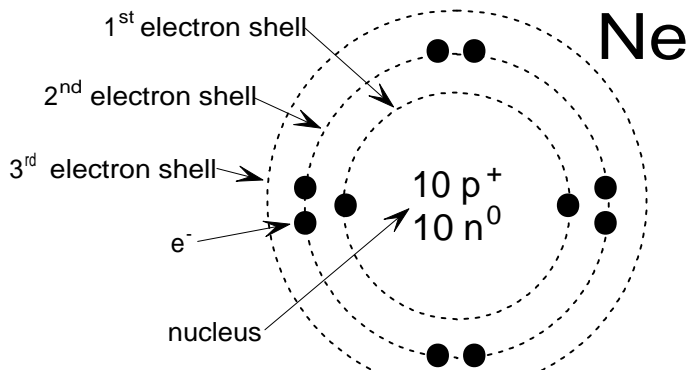
### Atomic Particles

Name	S	Mass	Charge
proton	p <sup>+</sup>	1	+1
neutron	n <sup>0</sup>	1	0
electron	e <sup>-</sup>	0	-1

### Guide to the Periodic Table

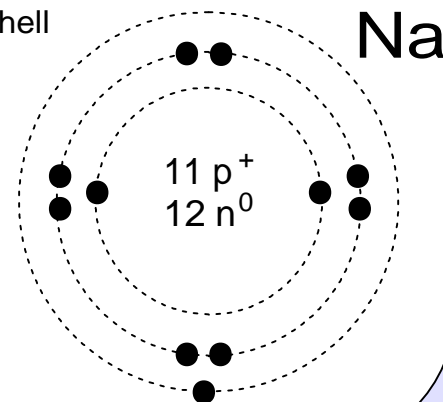
1. The periodic table is composed of element boxes. Each box contains an element's atomic number, atomic mass, name and symbol.
2. The element boxes are arranged in a table of 7 rows and 18 columns.
3. The columns are called groups and represent elements that have similar chemical properties due to the same number of electrons in their last shell.
4. The rows are called periods. Elements in a row are arranged by increasing atomic number.
5. The period number is the number of the electron shell that elements of that period are filling. Elements of a row have different chemical properties.

### Atomic Structure (Bohr) Diagrams



Electron shells are filled from 1<sup>st</sup> shell outwards.

Shell	Max. e <sup>-</sup>
1 <sup>st</sup>	2
2 <sup>nd</sup>	8
3 <sup>rd</sup>	18 (8)



### Group 1A

### First Three Rows of the Periodic Table (Main Group Elements)

### Group 8A

Period 1	1 H 1.0079 Hydrogen	2A	3A	4A	5A	6A	7A	2 He 4.0026 Helium
Period 2	3 Li 6.941 Lithium	4 Be 9.0122 Beryllium	5 B 10.81 Boron	6 C 12.011 Carbon	7 N 14.007 Nitrogen	8 O 15.999 Oxygen	9 F 18.998 Fluorine	10 Ne 20.179 Neon
Period 3	11 Na 22.990 Sodium	12 Mg 24.305 Magnesium	13 Al 26.982 Aluminum	14 Si 28.086 Silicon	15 P 30.974 Phosphorus	16 S 32.06 Sulfur	17 Cl 35.453 Chlorine	18 Ar 39.948 Argon

